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Physical chemical properties of matter worksheet

We are all surrounded by matter every day. Anything we use, touch, eat, etc. is an example of matter. Matter can be defined or described as anything that takes up space, and is composed of tiny particles called atoms. You must display the two mass and volume properties. Different types of matter can be distinguished through two components: composition and properties. The composition of matter refers to the different components of matter along with its relative proportions. The properties of matter refer to the qualities/attributes that distinguish one sample of matter from another. These properties are generally grouped into two categories: physical or chemical. Figure (PageIndex{1}): Organizational breakdown of the chemical and physical properties of matter. Physical properties can be observed or measured without changing the composition of matter. Physical properties are used to observe and describe matter. The physical properties of materials and systems are often described as intensive and extensive properties. This classification refers to the dependence of properties on the size or extent of the system or object in question. An intensive property is a bulk property, which means that it is a physical property of a system that does not depend on the size of the system or the amount of material in the system. Examples of intensive properties include the temperature, refractive index, density, and hardness of an object. When cutting a diamond, the pieces maintain their intrinsic hardness (until their size reaches a few atoms thick). In contrast, an extensive property is additive for independent subsystems that do not interact. The property is proportional to the material quantity in the system. Intensive properties: A physical property that will be the same regardless of the amount of matter. Density, pigment or shadow color: electricity to flow through the substance malleability: if a substance can be flattened brightness: how bright the substance looks Extensive Properties: A physical property that will change if the amount of matter changes. mass: How much matter in the sample volume: how much space the sample length takes: how long the sample is Change of physical change in which the physical appearance of the matter is modified, but the composition remains unchanged. A physical change takes place without any change in molecular composition. The same element or compound is present before and after the change. The same molecule is present through changes. Physical changes are related to physical properties, as some measurements require changes. The three main states of matter are: Solid, Liquid, Solid Gas is distinguished by a fixed structure. Its shape and volume do not change. In a solid, atoms are firmly packed together in a fixed arrangement. The liquid is distinguished by its malleable shape (it is able to form in the shape of its container), but constant volume. In a liquid, atoms are close, but not in a fixed arrangement. The gas is made up of separate atoms. However, unlike solid and liquid, a gas has no fixed shape and volume. Pageindex example (PageIndex{1}): Physical change When H₂O liquid water freezes in a solid state (ice), it appears changed; However, this change is only physical as the composition of the constituent molecules is the same: 11.19% hydrogen and 88.81% oxygen per mass. Pageindex figure(2)): Physical change: Ice merging is a physical change. Wikipedia. The chemical properties of matter describe its potential to undergo any chemical change or reaction by virtue of its composition. What elements, electrons and bonding are present to give the potential for chemical change. It is quite difficult to define a chemical property without using the word change. Eventually you should be able to look at the formula of a compound and the state of some chemical property. Right now this is very difficult to do and you are not expected to be able to do so. For example, hydrogen has the potential to ignite and explode under the right conditions. This is a chemical property. Metals generally have chemical properties of reacting with an acid. Zinc reacts with hydrochloric acid to produce hydrogen gas. This is a chemical property. Chemical change results in one or more substances of completely different composition from the original substances. Elements and/or compounds at the beginning of the reaction are rearranged into new compounds or product elements. A CHEMICAL CHANGE alters the composition of the original matter. Different elements or compounds are present at the end of the chemical change. The atoms in the compounds are reorganized to make new and different compounds. Example of corrosion[1] of metals is the unwanted oxidation of metals that results in metal oxides. [2 Mg + O₂ of right arrow figure 2 MgO] (PageIndex{3}): Chemical change: A burning magnesium tape with very short exposure for oxidation details. Captain John Yossarian (Wikipedia) The following questions are multiple choice. 1. Milk becomes bitter. This is a chemically owned chemically owned property of _____. None of the previous 2. HCl being a strong acid is a _____. The sawn wood in two is _____. Chemical Change, Physical Change Physical Change, Chemical Change Chemical Property Chemical, Physical Change Physical Property, Chemical Change None of the previous 3. CuSO₄ dissolves in water Chemical Change Chemical Property Physical Property None of the previous 4. Aluminum phosphate has a density of 2.566 g/cm³ Chemical Change Chemical Property Chemical Property Physical Property None of the previous 5. Which of the following are examples of subject matter? A dog carbon dioxide copper ice cubes (l) nitrate A mobile car 6. The formation of gas bubbles is a sign of exchange rate? 7. True or False: Increased bread is a physical property. 8. True or False: Potato removal is a Change. 9. Is it sunlight? 10. Lead mass is a _____ property. Chemical solutions change chemical property, physical change physical property All previous chemical False True No physical property References Petrucci, Bissonnette, Herring, Mature. General Chemistry: Modern Principles and Applications. Tenth ed. Upper Saddle River, NJ 07458: Pearson Education Inc., 2011. Cracolice, Peters. Fundamentals of Introductory Chemistry An Active Learning Approach. Second ed. Belmont, CA 94001:Brooks/Cole, 2007. Collaborators and Attributions We are all surrounded by matter on a daily basis. Anything we use, touch, eat, etc. is an example of matter. Matter can be defined or described as anything that takes up space, and is composed of tiny particles called atoms. 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